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Assignment 1

1.) Write any two properties of acid, base and salt.

→ Any two properties of acid, base & salt are:

Acid;
(i) Strong acids are corrosive in nature.
(ii) Acids dissolve in water and give hydrogen ions.

Base;
(i) Bases are soapy in touch and bitter in taste.
(ii) Strong bases burn our skin.

Salt;
(i) Some salts are salty in taste but most salts are bitter.
(ii) Salts conduct electricity in molten or solution state.

→

2. Differentiate between acid and base.

→

Acid	Base
- They give Hydrogen ions when dissolved in water.	- They give hydroxyl ions when dissolved in water.
- They are sour in taste.	- They are bitter in taste.

3. Make a chart to show the action of indicators [red and blue litmus paper, methyl orange, and phenolphthalein) in acidic, basic and salt solution.

→

S.N.	Indicators	Color in acid solution	Color in basic solution	Color in neutral salt solution.
1)	Red Litmus Paper	No change in color	Changes into blue	No changes in color
2)	Blue Litmus Paper	Changes into red	No changes in color	No change in color
3)	Methyl orange	Changes into red	Changes into yellow	No change in color
4)	Phenolphtalein	No change in color	Changes into pink	No change in color
5)	Red cabbage juice	Changes into color	Changes into green	Changes into rose red.

4. Write any two differences between bases & alkali.

→

Base	Alkali
- All metallic oxides are base.	- Only water soluble bases are alkalis.
- All bases are not alkalis.	- All alkalis are bases.

5. Write down the uses of the following.

a) Potassium Hydroxide (caustic potash)
→ Used in batteries.

b) Sodium hydroxide (caustic soda)
→ To make detergents, soap, etc.

c) Calcium Sulphat (Gypsum salt)
→ To plaster fractured bones.

d) Copper sulphate
→ To make fungicides

e) Magnesium sulphate
→ To prevent low blood magnesium.

f Sodium chloride
→ Used as preservative in foods.

5. pH scale
→ To measure the strength of acidic and basic solution.

6. Define universal indicator. How is it different from ordinary indicator?
→ Universal indicator is a special kind of indicator which is used to measure the strength of acidity or alkalinity.
→ It is different from ordinary indicators as ordinary indicators only show whether the substance is acid, base or salt but Universal indicator shows ~~the strength with~~ whether the substance is acid, base or salt with its strength.

7. All ~~bases~~ alkalis are ~~alkalis~~ bases but all bases are not alkalis. Explain.
→ ...because alkalis are water soluble bases but all bases are not water soluble.

8. Why acid should be handled with care?

→ ∴ because they are corrosive in nature.